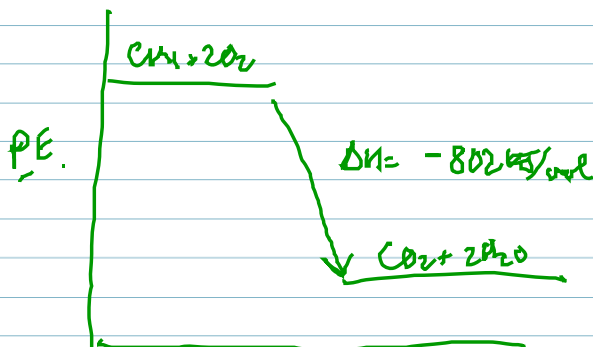
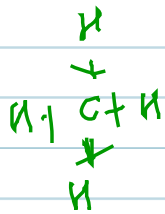
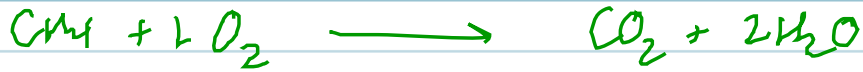


$$\Delta H_{\text{rxn}} \approx 2632 \text{ kJ/mol} + \ominus 3434 \text{ kJ/mol} = \ominus 802 \text{ kJ/mol}$$





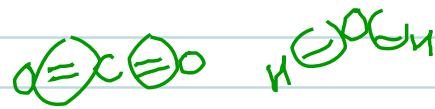
$$0 \neq 0$$

$$0 \neq 0$$

$$4(\text{C}-\text{H}) + 2(\text{O}=\text{O})$$

$$4(411) + 2(494)$$

$$2632 \text{ kJ/mol}$$

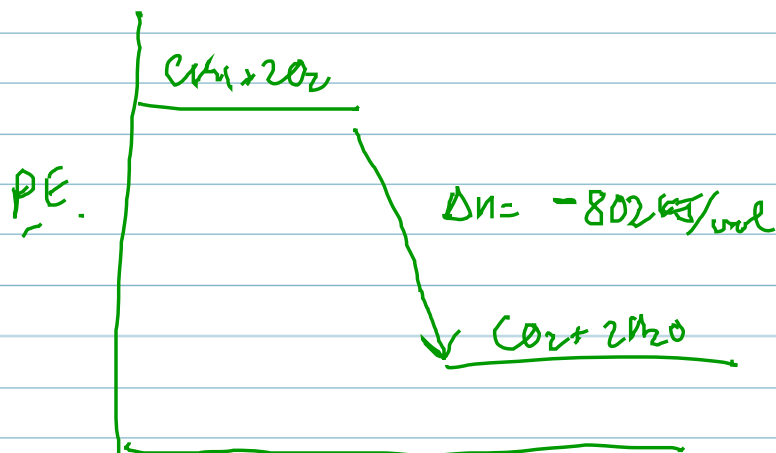


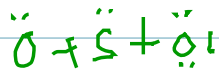
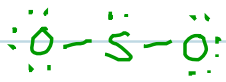
$$2(\text{C}=\text{O}) + 4(\text{O}-\text{H})$$

$$2(-799) + 4(-459)$$

$$-3434 \text{ kJ/mol}$$

$$\Delta H_{\text{rxn}} = 2632 \text{ kJ/mol} + (-3434 \text{ kJ/mol}) = -802 \text{ kJ/mol}$$

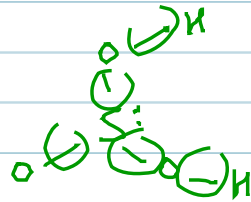
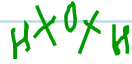




$$2(\text{S}=\text{O}) + 2(\text{O}-\text{H})$$

$$2(532) + 2(459)$$

$$\underline{1982 \text{ kJ/mol}}$$



$$3(\text{S}-\text{O}) + 2(\text{O}-\text{H})$$

$$3(-364) + 2(-459)$$

$$\underline{-2010 \text{ kJ}}$$

$$\Delta H = 1982 \text{ kJ/mol} + -2010 \text{ kJ/mol} = -28 \text{ kJ/mol}$$